

Sulfide-Arrested Growth of Gold Nanorods

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The growth of gold nanorods can be arrested at intermediate stages by treatment with Na₂S, providing greater control over their optical resonances. Nanorods prepared by the seeded reduction of AuCl₄ in aqueous cetyltrimethylammonium bromide solutions in the presence of AgNO₃ typically exhibit a gradual blueshift in longitudinal plasmon resonance, over a period of hours to days. This “optical drift” can be greatly reduced by adding millimolar concentrations of Na₂S to quench nanorod growth, with an optimized sulfur:metal ratio of 4:1. The sulfide-treated nanorods also experience a marked redshift as a function of Na₂S concentration to produce stable plasmon resonances well into the near-infrared. Sulfide treatment permitted a time-resolved analysis of nanorod growth by transmission electron microscopy, revealing two distinct periods: an initial growth burst ($t < 15$ min) that generates dumbbell-shaped nanorods with flared ends and a slower phase ($t > 30$ min) favoring growth around the midsection, leading to nanorods with the more familiar oblate geometry. The blueshift in plasmon resonance that accompanies the dumbbell-to-oblate shape transition correlates more strongly with changes in the length-to-midsection (L/D_1) ratio rather than the length-to-end width (L/D_2) ratio, based on the empirical relationship introduced by El-Sayed and co-workers (Link, S.; Mohamed, M. B.; El-Sayed, M. A. *J. Phys. Chem. B* **1999**, *103*, 3073–3077).

Introduction

Gold nanostructures offer excellent prospects for probing biological systems. The base material is biochemically inert and can be manipulated at nanometer length scales to support localized surface plasmon resonances at visible and near-infrared (NIR) wavelengths. The latter is well-known to penetrate biological tissues with relatively high transmittivities: wavelengths shorter than 800 nm are absorbed by hemoglobin and melanin, and wavelengths greater than 1300 nm are strongly attenuated by water. Plasmonic nanomaterials with NIR resonances are currently being investigated as biomolecular sensors,^{1–3} as contrast agents for biomedical imaging,⁴ and as mediators of photodynamic therapy.⁵

Gold nanorods are particularly well-suited for biosensing and biomedical imaging at NIR frequencies because their plasmon resonance modes can be tuned as a function of aspect ratio.^{6–9} Nanorods have been synthesized in the

presence of micelle-forming surfactants such as cetyltrimethylammonium bromide (CTAB) by electrolysis at cathodic potentials¹⁰ or by the chemical reduction of AuCl₄ ions.^{11–15} With regard to the latter, several groups have shown that seeded growth under mildly reducing conditions can produce nanorods with exquisite control over size and shape, as demonstrated by transmission electron microscopy (TEM).^{11–13} Sau and Murphy have recently demonstrated that nanorods with aspect ratios of 2–4 can be formed in nearly quantitative yields when AgNO₃ is used as an additive; in some cases, a dumbbell-like morphology appears to be favored over the cylindrically symmetric oblate geometry.^{11c} Dumbbell-shaped nanostructures have also been reported by Liu and Guyot-Sionnest¹⁶ and by Chang and co-workers,¹⁷ who used

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performed gold nanorods to nucleate the reduction of silver ions. In contrast, a study by Mulvaney and co-workers did not reveal any dumbbell-shaped particles when nanorods were prepared in the absence of AgNO_3 .¹³

The seeded growth of Au nanorods in the presence of AgNO_3 promises to be a practical synthesis with respect to both yield and uniformity. However, we have observed that the plasmon resonances of such nanorods are often unstable and can drift toward shorter wavelengths by as much as 100 nm over a period of hours to days. Applications involving plasmon-enhanced spectroscopies require substrates with stable, well-defined resonances and line shapes: a shift of a few nanometers can reduce signal intensities by orders of magnitude, particularly if nonlinear optical enhancements are involved.^{9,18,19} The nanorods' optical instability thus becomes an important issue to address, as it limits them from reaching their full technological potential.

Here we demonstrate that the optical drift in Au nanorods can be greatly reduced by the controlled addition of Na_2S . Sulfide ions have a strong affinity for heavy metals in their ionic and metallic states and are commonly used in ore beneficiation and metal surface passivation.²⁰ The sulfide treatment provides an additional benefit by modulating the dielectric function at the nanorod surface, shifting the plasmon resonances toward longer wavelengths. The sulfide-treated nanorods are stable at room temperature after isolation from the reaction mixture and can be stored for months without degradation. Finally, the sulfide treatment permits us to obtain reliable insights on nanorod growth after nucleation, enabling a correlation between nanostructure and optical resonance.

Experimental Section

Reagents. Cetyltrimethylammonium bromide (CTAB), AgNO_3 , and HAuCl_4 were purchased from Alfa Aesar and used as supplied. The quality of the CTAB is particularly important as minor impurities (most likely cetyldimethylamine) can result in much faster rates of reduction, with subsequently poor control over the reaction kinetics. Deionized water was obtained using an ultrafiltration system (Milli-Q, Millipore) with a measured resistivity above $18 \text{ M}\Omega\cdot\text{cm}$ and passed through a $0.22\text{-}\mu\text{m}$ filter to remove particulate matter.

Au Nanorod Synthesis. Growth and seed solutions were prepared according to the protocol described by Nikoobakht and El-Sayed.¹² Au seeds ($d_{\text{av}} = 3.5 \text{ nm}$) were prepared by mixing equal portions of an aqueous 0.2 M CTAB solution (5 mL) and a 0.5 mM HAuCl_4 solution (5 mL) at room temperature in a clean glass vial, followed by an injection of a cold ($0 \text{ }^\circ\text{C}$) 10 mM NaBH_4 solution (0.6 mL). The reaction mixture was stirred for 2 min with a magnetic stir bar prior to its addition to the growth solution.

The nanorod growth solution was prepared by combining an aqueous solution of 0.2 M CTAB (5 mL) with 4 mM AgNO_3 (0.24 mL) in a 40-mL vial, followed by immersion in an ultrasonic

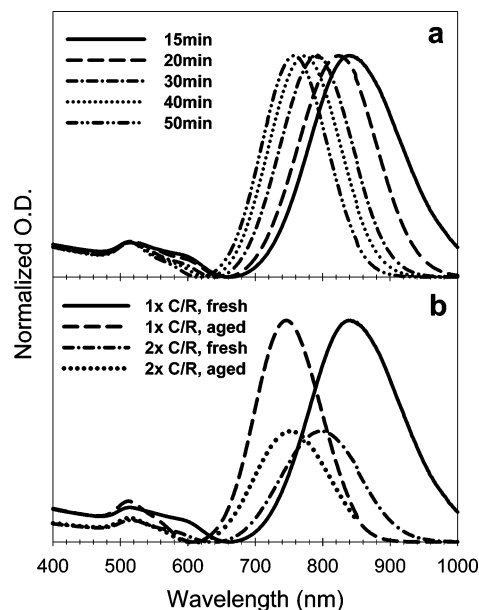


Figure 1. Normalized absorbance spectra of Au nanorods subjected to centrifugation and redispersion (C/R) in deionized water. (a) Nanorods subjected to C/R at various intervals after seed addition ($t = 15, 20, 30, 40,$ and 50 min). (b) Nanorods ($t = 25 \text{ min}$) after one or two C/R cycles, with spectra taken immediately after C/R (fresh) and several days afterward (aged).

cleaning bath (Branson, 42 kHz) for 10 min. The growth solution was then treated with 1 mM HAuCl_4 (5 mL) and 78.8 mM ascorbic acid ($70 \text{ }\mu\text{L}$), followed by an injection of the seed solution ($12 \text{ }\mu\text{L}$; see above). The reaction mixture was stirred by hand for 30 s and then allowed to sit at room temperature for a designated period of time (t) before being quenched with 15 mL of Na_2S solution at various concentrations, with the mole ratio of S to metal (Au and Ag) ranging from 1:2 (0.125 mM) to 4:1 (1 mM). The sulfide-treated nanorods were transferred to plastic centrifuge tubes and subjected to centrifugation at $12500g$ (1 min at maximum speed), then decanted from the supernatant and redispersed in deionized water ($6\text{--}8 \text{ mL}$), and characterized by optical absorption spectroscopy using a Cary-50 visible-NIR spectrophotometer (Varian) with a 1-cm cell path length.

Results and Discussion

In the course of a typical nanorod synthesis, plasmon resonance peaks appeared within 15–20 min in the NIR and reached their maximum absorbance intensities after 50–60 min accompanied by a blueshift of the longitudinal plasmon peak toward the visible, in accord with the recent report of Sau and Murphy (see Figure 1a).^{11c} It must be mentioned that the resonance wavelengths produced after a given time interval varies from batch to batch and appear to be quite sensitive to nucleation conditions. We have observed that nanorod batches prepared and isolated under apparently identical conditions can differ in their optical resonances by 5–30 nm, precluding an exact comparative analysis for different reaction conditions. Nonetheless, the same trends are always observed with respect to the blueshift during the course of the reaction, as well as the subsequent optical drift over time. Therefore, all discussions concerning the optical stability of nanorods are based on post-synthetic changes in plasmon resonance, relative to aliquots of freshly prepared, untreated nanorods.

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Table 1. Particle Size Analysis of Au Nanorods Using Various Sampling Conditions^a

sample	count (N)	length L (nm)	midsection D ₁ (nm)	end width D ₂ (nm)	width ratio (D ₁ /D ₂)	aspect ratio (L/D ₂)
A (no C/R)	100	57.7 ± 8.2	19.1 ± 4.2	22.4 ± 3.8	0.86 ± 0.19	2.62 ± 0.39
B (1 × C/R)	100	48.3 ± 5.6	14.7 ± 2.0	16.3 ± 2.3	0.91 ± 0.05	2.99 ± 0.37
C (1 × C/R, 24 h later)	118	46.6 ± 11	15.0 ± 1.6	15.3 ± 1.9	0.96 ± 0.05	3.00 ± 0.48

^a Nanorods isolated 30 min after seed addition; see main text for sample preparation conditions. Errors represent one standard deviation.

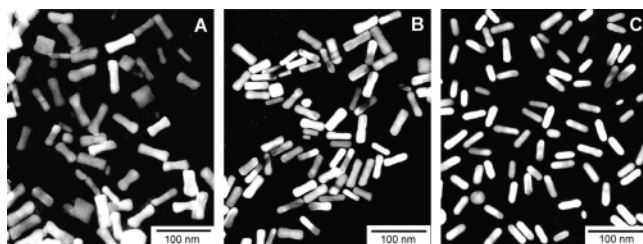


Figure 2. TEM images (Philips CM-10, 80 kV) of Au nanorods isolated 30 min after seed addition. Nanorod suspensions were cast onto carbon-coated Cu grids, then blotted after 30 s and dried in air. Sample A: Nanorods deposited directly from the reaction mixture. Sample B: Nanorods subjected to C/R in deionized water, just prior to sample preparation. Sample C: Same batch of nanorods as in B, aged 24 h prior to sample preparation.

Our initial attempts to isolate nanorods with specific optical resonances involved high-speed centrifugation, followed by redispersion of the precipitants in deionized water.¹² This proved not to be effective: the optical absorption of the recovered nanorods continued to drift toward shorter wavelengths over a period of several days, even after two consecutive cycles of centrifugation/redispersion (C/R; see Figure 1b). TEM analysis further revealed that the conditions and timing of sample preparation had a strong influence over nanorod size and shape: particles cast directly from the seeded solutions onto TEM substrates (sample A) exhibited dumbbell-like morphologies, whereas those collected after one C/R cycle (samples B and C) underwent a gradual structural transition to a cylindrically symmetric form (see Figure 2). A likely explanation for this evolution is that the unreacted Au ions were not freely suspended in solution but were adsorbed instead onto the surfaces of CTAB micelles, which co-precipitated with the nanorods upon centrifugation and could be subsequently redispersed. The virtual absence of free Au ions in the presence of CTAB and the very low solubility of the Au–CTAB complex have been confirmed in previous studies.²¹ Mulvaney and co-workers have recently provided strong evidence for a nanorod elongation mechanism that involves a collision-activated process with Au–micelle complexes.¹³

Particle size analysis of samples A–C provided some additional insights in the changes of nanorod structure as a function of sample preparation (see Table 1). Two trends are immediately apparent: (i) the nanorod lengths (L) and end widths (D_2) decrease from samples A–C, and (ii) the midsection-to-end width ratios (D_1/D_2) increase from A–C. The first trend indicates a decrease in residual metal ion, which contribute to nanorod growth even while drying on the TEM substrates, whereas the second trend suggests a late growth stage around the nanorod midsection. The D_1/D_2 ratios may be overestimated by an acceleration in lateral (D_1) growth during TEM sample preparation: as the wetting

layer evaporates, dewetting forces can concentrate residual solute into regions of negative curvature, promoting localized reduction of Au(I) ions. This lack of synthetic control underscores the difficulty of characterizing the nanorods in the presence of unreacted Au ions.

Treating the reaction mixtures with sodium sulfide was found to be an effective means of quenching nanorod growth. Reactions were set up in parallel and treated with Na₂S solutions at $t = 15, 30,$ and 50 min after addition of the seed solution, followed 15 min later by centrifugation and redispersion of the nanorods in deionized water. The mole ratio of S to the total metal content (Au and Ag, or M) was varied from 1:2 to 4:1. As before, a blueshift in λ_{\max} was observed during the course of nanorod growth, but this was arrested immediately upon addition of Na₂S.

Monitoring the sulfide-treated nanorods over a 6-day period revealed large improvements in optical stability as a function of Na₂S addition, the effect being most pronounced for nanorods treated at an early growth stage (see Figure 3). Minimal drift was observed at a S:M ratio of 4:1. For nanorods quenched at $t = 15$ min, the optical drift over a 6-day period was reduced from -100 nm (no sulfide treatment) to -5 nm; for nanorods quenched at $t = 50$ min, the optical drift could be completely arrested (see Table 2). Further increases in sulfide concentration did not have any benefits; in fact, excessive sulfide led to the slow disintegration of the nanorods. The reduction in optical drift with Na₂S concentration suggests that sulfide serves primarily as a scavenger for Au and Ag ions, effectively removing them from the reaction mixture. Zhang and co-workers have provided evidence that AuCl₄ ions are reduced by Na₂S to form metallic Au nanoparticles rather than Au₂S.²²

Treating the Au nanorods with Na₂S also induced a concentration-dependent increase in plasmon resonance toward longer wavelengths. This redshift, which we ascribe to modulations in the nanorods' dielectric function by sulfide adsorption, is again most pronounced for nanorods quenched at an early stage of growth (see Table 2). In contrast to the slow optical drift described previously, the sulfide-induced redshift is essentially a singular event. To confirm the influence of sulfide adsorption on plasmon resonance, previously untreated nanorod dispersions (aged for several days) were treated with Na₂S and examined by absorption spectroscopy. A concentration-dependent effect on the longitudinal plasmon resonance was again observed, with redshifts up to 29 nm for a S:M ratio of 4:1 (see Figure 4). It is interesting to note that a redshift was not induced at the lowest S:M ratio (1:2), implying that the sulfide did not reach

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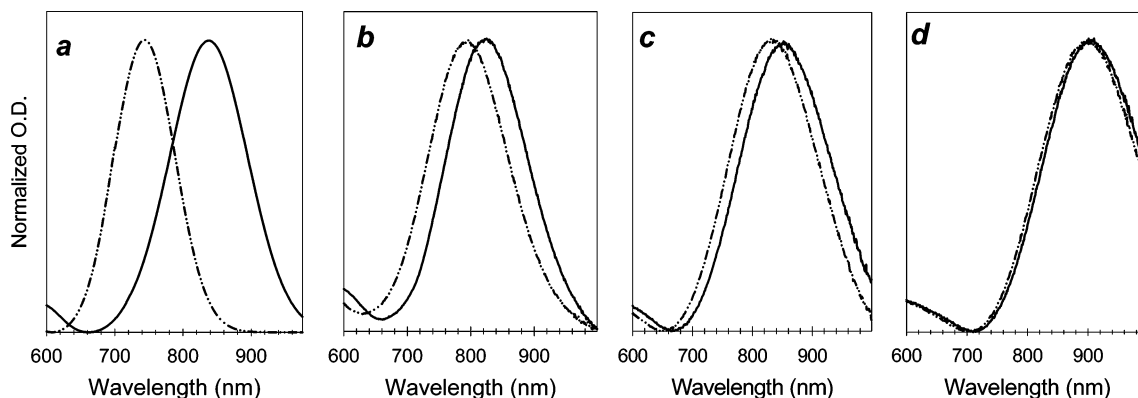


Figure 3. Time-dependent changes in the plasmon resonance of sulfide-treated Au nanorods, quenched at $t = 15$ min. Absorption spectra of nanorods were obtained immediately after one C/R cycle in deionized water (solid line) and after 6 days (dashed line). (a) no sulfide addition; (b) S:M = 1:2; (c) S:M = 1:1; (d) S:M = 4:1.

Table 2. Effect of Na_2S Treatment on Nanorod Plasmon Resonance and Optical Drift^a

sample		sulfur:metal ratio ^b			
		no S	1:2	1:1	4:1
$t = 15$ min	λ_{max}^c	840	824	851	905
	$\Delta\lambda^d$	-100	-31	-18	-5
$t = 30$ min	λ_{max}^c	792	771	789	807
	$\Delta\lambda^d$	-51	-11	-9	-5
$t = 50$ min	λ_{max}^c	758	756	768	783
	$\Delta\lambda^d$	-56	-11	-5	~ 0

^a All nanorods were isolated immediately after Na_2S treatment by centrifugation and redispersed in deionized water. ^b Metal includes total Au and Ag ions added to initial reaction mixture. ^c Taken immediately after 1 C/R cycle, in nm. ^d Shift in λ after 6 days, in nm. Differences in initial wavelengths prior to sulfide treatment preclude a fully quantitative comparison between samples.

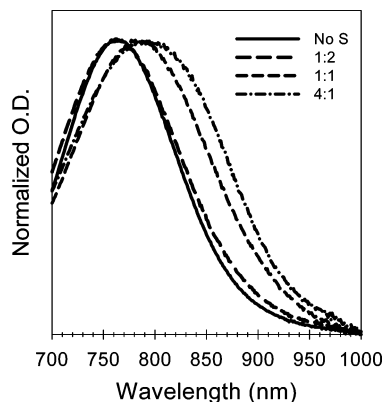


Figure 4. Sulfide-induced shifts in plasmon resonance on aged nanorods. Nanorods were subjected to one C/R cycle at $t = 20$ min and aged for several days prior to Na_2S addition. Shifts in plasmon resonance (measured immediately after sulfide addition) are as follows: S:M = 1:2, +0 nm; S:M = 1:1, +20 nm; S:M = 4:1, +29 nm.

the nanorod surface. Changes in plasmon resonance as a function of the medium dielectric are well-established for Au nanospheres²³ as well as for anisotropic nanoparticles such as nanotriangles and nanorods,²⁴ but the latter are the most sensitive to changes in their local environment.²⁵

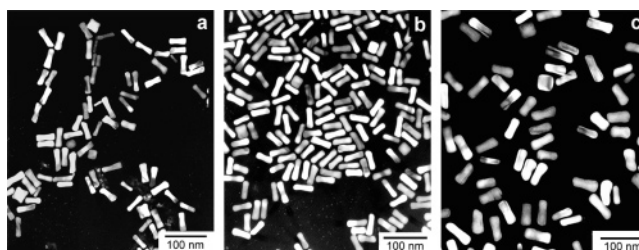


Figure 5. TEM images (Philips CM-10, 80 kV) of Au nanorods after Na_2S treatment. Nanorod suspensions were subjected to one C/R cycle prior to TEM sample preparation. Reaction times before quenching with 0.2 mM Na_2S : (a) $t = 15$ min; (b) $t = 30$ min; (c) $t = 50$ min.

TEM analysis of the sulfide-treated Au nanorod suspensions provided an unambiguous record of the time-dependent changes in nanorod morphology over the course of the reaction (see Figure 5 and Table 3; additional images in Supporting Information). Nanorods were treated at a S:M ratio of 1:2 at several different time intervals ($t = 15, 20, 30, 40,$ and 50 min), then immediately isolated and deposited onto Formvar-coated Cu grids. These samples differed significantly from the untreated nanorod samples prepared under similar conditions, with respect to changes in size and aspect ratios (cf. Figure 2 and Table 1).

The TEM data reveals an initial growth burst predominantly in the longitudinal direction to produce dumbbell-shaped nanorods ($t < 15$ min), followed by an intermediate period of monotonic growth in both longitudinal (L) and lateral (D_1 and D_2) directions ($t = 15$ – 30 min). The flaring in the dumbbell-shaped nanorods is most pronounced in the first 30 min after seed addition, with an approximately 20% difference between D_1 and D_2 . A new phase appears after $t > 30$ min, in which lateral growth around the nanorod midsection is faster than at the ends: while the growth rate remains unchanged in the L and D_2 dimensions, growth in the D_1 dimension increases abruptly. The shift in growth kinetics to the D_1 dimension is presumably responsible for producing the final oblate geometry (cf. Figure 2c) and subsequent blueshift in plasmon resonance. Two additional observations are worth mentioning: (i) Sulfide-treated nanorods with dumbbell-like morphologies did not lose their shape over time and were thus not susceptible to CTAB-mediated reshaping.²⁶ (ii) Untreated nanorods isolated after $t = 50$ min were not uniform in shape but observed to be a mixture of dumbbell and oblate geometries, indicating

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Table 3. Size Analysis of Sulfide-Treated Nanorods^a

<i>t</i> (min)	count (<i>N</i>)	length <i>L</i> (nm)	midsection <i>D</i> ₁ (nm)	end width <i>D</i> ₂ (nm)	width ratio (<i>D</i> ₁ / <i>D</i> ₂)	aspect ratio (<i>L</i> / <i>D</i> ₁)	aspect ratio (<i>L</i> / <i>D</i> ₂)
15	95	44.3 ± 13	11.5 ± 1.5	14.5 ± 1.9	0.79 ± 0.06	3.99 ± 0.83	3.17 ± 0.62
20	103	48.9 ± 7.7	13.2 ± 1.5	16.9 ± 1.6	0.80 ± 0.05	3.71 ± 0.71	2.98 ± 0.47
30	117	50.5 ± 7.5	13.7 ± 1.6	17.2 ± 1.9	0.80 ± 0.04	3.71 ± 0.63	2.96 ± 0.51
40	91	55.2 ± 11.0	18.6 ± 2.5	21.1 ± 3.2	0.89 ± 0.05	2.96 ± 0.65	2.63 ± 0.59
50	103	58.1 ± 6.3	21.4 ± 2.1	24.2 ± 2.5	0.89 ± 0.05	2.77 ± 0.41	2.45 ± 0.37

^a Nanorods were isolated by one C/R cycle in deionized water, prior to TEM sample preparation. Errors represent one standard deviation.

Table 4. Longitudinal Plasmon Resonance Analysis of Dumbbell-shaped Nanorods^a

<i>t</i> (min)	λ_{\max} , expt (nm) ^b	λ_{\max} , calcd using <i>L</i> / <i>D</i> ₁ ^c	λ_{\max} , calcd using <i>L</i> / <i>D</i> ₂ ^c
15	824	806	721
30	771	777	699
50	756	679	646

^a Nanorods were quenched with Na₂S and isolated by one C/R cycle in deionized water. ^b Experimental λ_{\max} values taken from Table 2 using S:M = 1:2. ^c λ_{\max} calculated using eqs 1 and 2 and the values for *L*/*D*₁ and *L*/*D*₂ from Table 3.

variable rates of growth in the *D*₁ direction (see Supporting Information). The latter observation suggests that nanorods with the highest uniformity can be produced by quenching their growth at an early stage, within 30 min of their nucleation.

The presence of two different lateral dimensions *D*₁ and *D*₂ raises the interesting question whether either of these would be useful for characterizing the nanorods' longitudinal plasmon resonances. Link et al.²⁴ derived a parameterized equation based on Mie–Gans theory to establish a linear correlation between the aspect ratio of cylindrical nanorods and their plasmon resonance wavelengths:

$$\lambda_{\max} = (33.34R - 46.31)\epsilon_m + 472.31 \quad (1)$$

$$\epsilon_m = (103.70R - 80.20)/(33.34R - 46.31) \quad (2)$$

where *R* is the nanorod aspect ratio and ϵ_m is the effective dielectric medium of the surrounding micellar layer, which also varies as a function of *R*.^{24,27} The application of eqs 1 and 2 is only useful to the first degree of approximation, as the nanorod shapes and micellar materials used by Link et al.²⁴ are different from those in our study. Further modulation in the local dielectric medium by sulfide adsorption is also possible, although this effect appears to be negligible for a S:M ratio of 1:2 (cf. Table 2 and Figure 4). Nevertheless, the comparison is sufficient to demonstrate that the aspect ratio *L*/*D*₁ provides a better approximation of the longitudinal plasmon resonance wavelength than *L*/*D*₂ (see Table 4).

With respect to the growth mechanism of the dumbbell-shaped nanorods, recent evidence from the Murphy and Chang groups have shown that the {111} facets at the nanorod ends are receptive to Ag ion deposition.^{11c,17} In



Figure 6. TEM image (Philips CM-10, 80 kV) of anisotropic Au nanoparticles prepared in the absence of AgNO₃, followed by quenching with Na₂S (*t* = 25 min). Nanorod suspensions were subjected to one C/R cycle prior to sample preparation.

particular, the latter group has established that Ag metal can be deposited onto the nanorod ends by adjusting the reducing power of the growth solution. To test the role of AgNO₃ in the growth of dumbbell-shaped particles, nanorods were prepared by seeded growth in the absence of AgNO₃ and then quenched with Na₂S at various time intervals. The TEM images revealed several types of anisotropic nanoparticles including rods, triangles, and hexagonal and trapezoidal prisms, all of which appeared to have well-defined planar facets; however, no dumbbell-shaped particles could be observed (see Figure 6). This confirms that Ag plays a key role in dumbbell formation during the initial stages of nanorod growth. Although the precise mechanism remains to be determined, one possibility may involve the reduction of Ag ions by unreacted borohydride in the seed solution.²⁸

Conclusion

The CTAB-mediated growth of Au nanorods can be arrested by adding sulfide at millimolar concentrations, resulting in chemically and optically stable species with specific NIR resonances. The stabilizing effect of sulfide can be attributed to two processes: (i) the reduction or quenching of unreacted metal salts and (ii) the adsorption of sulfide onto the nanorod surfaces. The latter has the added benefit of shifting the plasmon resonance further into the NIR spectral region. The structural and optical stability of the sulfide-treated nanorods make them attractive candidates as contrast agents for optical coherence tomography and other biophotonic applications.

(26) Aguirre, C. M.; Kaspar, T. R.; Radloff, C.; Halas, N. J. *Nano Lett.* **2003**, *3*, 1707–1711.

(27) Revisions to eqs 1 and 2 have been suggested by Wang and co-workers, using a smaller effective value for ϵ_m (see: Yan, B.; Yang, Y.; Wang, Y. J. *J. Phys. Chem. B* **2003**, *107*, 9159). However, λ_{\max} values calculated using their revised formulation are significantly underestimated for aspect ratios greater than 3, when compared with the experimental values.

(28) Mulvaney and co-workers have commented on the importance of decomposing excess borohydride to avoid inadvertent nucleation; see ref 13.

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Supporting Information Available: Additional TEM images of Au nanorods before and after treatment with 0.2 mM Na₂S. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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